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the chemical properties of that element. Atoms cannot be broken down into smaller particles that will have the same chemical and physical properties as the atom. A single atom is so small that its size is difficult to imagine.

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hydrogen in  $MgH_2$ . 0 c. sulfur.

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c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol (b) divided by the mass of 1 mol  $nRT$  (c) multiplied by 22.4 L (d) divided by 22.4 L

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For the expression  $V =$  (a) increasing P  
(b) decreasing T

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